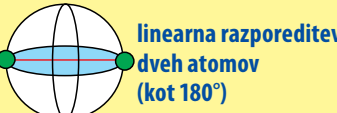
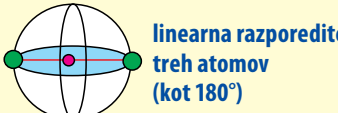




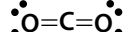

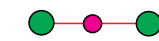

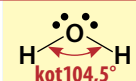
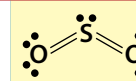
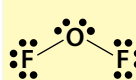
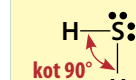


F ¹⁻ fluorid	Br ¹⁻ bromid	O ²⁻ oksid	N ³⁻ nitrid	C ⁴⁻ karbid
Cl ¹⁻ klorid	I ¹⁻ jodid	S ²⁻ sulfid	P ³⁻ fosfid	H ¹⁻ hidrid
1 mono	3 tri	5 penta	7 hepta	9 nona
2 di	4 tetra	6 hekso	8 okta	10 deka

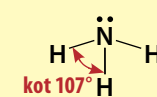
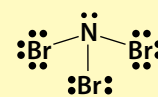
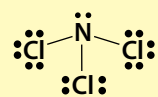
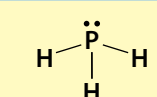
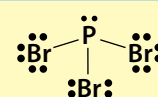
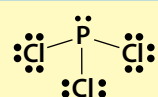
Oblike simetričnih molekul

 <p>linearna razporeditev dveh atomov (kot 180°) H₂, N₂, O₂, F₂, Cl₂, Br₂, I₂</p>	 <p>linearna razporeditev treh atomov (kot 180°) BeH₂, BeX₂: (BeF₂, BeCl₂, BeBr₂, BeI₂)</p>
 <p>planarna razporeditev treh atomov (kot 120°) BH₃, BX₃: (BF₃, BCl₃, BBr₃, BI₃)</p>	 <p>tetraedrična razporeditev štirih atomov (kot 109,5°) SiH₄, SiX₄: (SiF₄, SiCl₄ ...)</p>
 <p>trikotno bipiramidalna razporeditev petih atomov (koti 90° in 120°) PH₅, PX₅: (PF₅, PCl₅, PBr₅, PI₅)</p>	 <p>oktaedrična razporeditev šestih atomov (vsi koti po 90°) SH₆, SX₆: (SF₆, SCl₆, SBr₆, SI₆)</p>
CO ₂  linearna oblika	CS ₂  linearna oblika

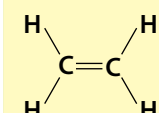
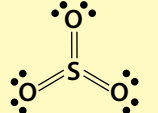
Oblike nesimetričnih molekul

Linearna oblika molekul		
HCN	H-C≡N: linearna oblika	C ₂ H ₂
	H-C≡C-H linearna oblika	
Kotna oblika molekul		
		
H ₂ O	 kotna oblika kot 104,5°	SO ₂
	 kotna oblika	
F ₂ O	 kotna oblika	H ₂ S
	 kotna oblika kot 90°	

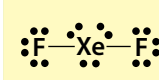
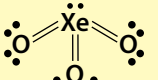
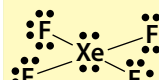
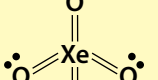
Piramidalna oblika molekul

NH ₃	NX ₃ (NF ₃ , NCl ₃ , NBr ₃ , NI ₃)	
 kot 107° H		
piramidalna oblika	piramidalna oblika	piramidalna oblika
PH ₃	PX ₃ (PF ₃ , PCl ₃ , PBr ₃ , PI ₃)	
		
piramidalna oblika	piramidalna oblika	piramidalna oblika

Planarne oblike molekul

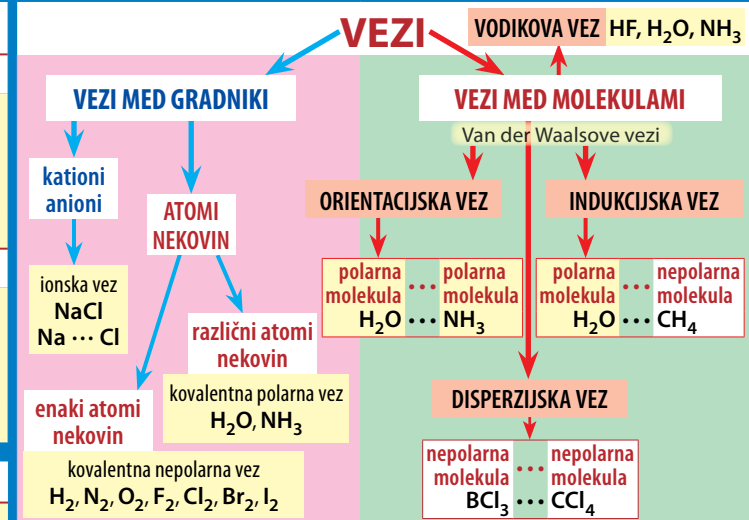
C ₂ H ₄	 planarna oblika	SO ₃	 planarna oblika
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Oblike molekul ksenona (Xe)

XeF ₂	 linearna oblika	XeO ₃	 piramidalna oblika
XeF ₄	 planarna oblika	XeO ₄	 tetraedrična oblika

Nepolarne molekule

H ₂ , N ₂ , O ₂ , F ₂ , Cl ₂ , Br ₂ , I ₂	BH ₃ , BX ₃ : (BF ₃ , BCl ₃ , BBr ₃ , BI ₃)
BeH ₂ , BeX ₂ : (BeF ₂ , BeCl ₂ , BeBr ₂ , BeI ₂)	C ₂ H ₂ , C ₂ H ₄ , CO ₂ , SO ₃
BH ₃ , BX ₃ : (BF ₃ , BCl ₃ , BBr ₃ , BI ₃)	SH ₆ , SX ₆ : (SF ₆ , SCl ₆ , SBr ₆ , SI ₆)
CH ₄ , CX ₄ : (CF ₄ , CCl ₄ , CBr ₄ , CI ₄)	He, Ne, Ar, Kr, Xe
SiH ₄ , SiX ₄ : (SiF ₄ , SiCl ₄ , SiBr ₄ , SiI ₄)	



Množina snovi - n, število delcev - N

$n = \frac{m}{M}$	$m = n \cdot M$	$M = \frac{m}{n}$	n - množina snovi [mol]
$n = \frac{N}{N_A}$	$N = n \cdot N_A$	$N_A = \frac{N}{n}$	m - masa snovi [g]
$\frac{m}{M} = \frac{N}{N_A}$	$m = \frac{M \cdot N}{N_A}$	$N = \frac{m \cdot N_A}{M}$	M - molska masa [g/mol]
			N - število delcev [atomi, ioni, molekule]
			N_A - Avogadrova konstanta [6,02 · 10 ²³ /mol]

Plinska enačba

$P \cdot V = n \cdot R \cdot T$	n - množina snovi [mol]
	P - tlak [kPa]
	V - volumen plina [L]
	M - molska masa snovi [g/mol]
	R - plinska konstanta [8,31 kPa L/mol K]
	T - temperatura [K]
$P \cdot V \cdot M = m \cdot R \cdot T$	
$T[K] = T[°C] + 273$	

Gostota plina - ρ

$\rho = \frac{P \cdot M}{R \cdot T}$	ρ - gostota plina [g/L]	P - tlak [kPa]
	R - plinska konstanta [8,31 kPa L/mol K]	T - temperatura [K]
$\rho = \frac{M}{V_m}$	V_m - molni volumen plina [L/mol]	
	M - molska masa [g/mol]	

Molski volumen plina - V_m

$V_m = \frac{R \cdot T}{P}$	n - množina plina [mol]	R - plinska konstanta [8,31 kPa L/mol K]
	P - tlak [kPa]	
	V - volumen plina [L]	T - temperatura [K]
$n = \frac{V}{V_m}$	V_m - molni volumen plina [L/mol]	

Kristal	Primeri	Gradniki	Vezi	Tališče	Električna prevodnost		
					Trdno	Talina	Raztopina
ionski	NaCl, CaCl ₂	kationi anioni	ionska	visoko	NE	DA	DA
kovalentni diamant	diamant	atomi ogljika	kovalentna nepolarna	visoko	NE	NE	NE
kovalentni grafit	grafit	atomi ogljika	kovalentna nepolarna	visoko	DA	NE	NE

Kristal	Primeri	Gradniki	Vezi	Tališče	Električna prevodnost		
					Trdno	Talina	Raztopina
kovalentni	SiO ₂ - kremen SiC - karborund	atomi nekovin	kovalentna polarna	visoko	NE	NE	NE
kovinski	Au, Fe, Zn	atomi kovin	kovinska	visoko nizko	DA	DA	—
molekulski	I ₂ , CO ₂ , H ₂ O, C ₁₁ H ₂₂ O ₁₁	molekule	molekulska	nizko	NE	NE	NE